**RATES OF REACTIONS**

**6.1 Reaction Rates Study Guide:**

Define (p. 346), reaction rate,

average reaction rate (p. 347)

Apply and Do, p. 350, # 1.

Determining the average reaction rate using graphical data,

Apply and Do, p. 352, # 1.

Instantaneous Rate of Reaction

Apply and Do, p. 356, #1 and 2

Apply and Do Review 6.1, # 1 and 2.

**6.2 Factors Affecting Reaction Rates**

Summarize and draw a small diagram to show why a reaction occurs:

p. 362, Chemical Nature of the reactants

p. 363, Concentration of reactants

p. 363, Surface Area

p. 363, Temperature of the reaction

p. 364, Presence of a catalyst,

Define, p. 364, biological catalyst, heterogeneous homogeneous catalyst.

Apply and do, p. 365, Review 6.2, # 1 and 2.

**6.3 Explaining Reaction Rates: p. 366.**

Do p. 372, #2 abc

Below Corresponds to Temperature of the Reaction System, 6.3 Explaining Reaction Rates

Apply and Do: p. 372, # 4-6.

6.3 Study Guide-

Define: Collision Theory , activation energy, temperature of the reaction system and what Maxwell-Boltzmann Distribution is and how it applies to Two Temperatures.

Chemical Nature of Reactants, How much energy does it take to break a single versus a double bond.(p. 370). ?

Concentration and surface area (p. 370)

Define and explain catalyst theory (p. 371)

**6.5 Study Guide-**

Define (p. 375): Rate Law Equation , Rate constant

Order of reaction

Total order of reaction

Define and understand (p. 376) initial rate and series of kinetic experiments.

Do and apply (p. 380), Practice #1-3

Do and apply (p. 382), Review, #1-2.